

## Homework Set for Session #12 (resists)

- 10-1. Read the attached article [Brunner, T., "Why optical lithography will live forever," J. Vac. Sci. Tech. B **21** pp. 2632-7 (2003)].
- What is the minimum value of  $k_1$  used in equation 1 for an optical system operating in a vacuum or in air.
  - Write a short paragraph on what you see as the major problem facing lithography for semiconductor manufacturing in the sub-100-nm domain over the next decade.

- 10-2. We will do a very rough comparison of photographic and photoresist sensitivity:

Assume our simple model for photographic film: that it consists of cubic blocks of AgBr grains tightly packed, that only one photon is needed to expose a block, and that only one block per vertical column need be exposed. Assume  $1 \times 1 \times 1 \mu\text{m}$  grains, and that an incident photon has a 99% probability of being absorbed as it passes through. How many ergs/cm<sup>2</sup> are needed for exposure at  $\lambda = 400 \text{ nm}$ ? How many would be needed if the grain size was  $0.2 \times 0.2 \times 0.2 \mu\text{m}$  instead of  $1 \times 1 \times 1 \mu\text{m}$ ?

Assume a polymer resist has a molecule size of  $1 \times 1 \times 1 \text{ nm}$ . Assume that full exposure corresponds to about 50% of the molecules absorbing a photon. Assume an absorption length of  $1 \mu\text{m}$  (i.e.,  $I = I_0 e^{-\alpha x}$ , where  $\alpha = 1 \mu\text{m}^{-1}$ ). How many ergs/cm<sup>2</sup> are required to "fully expose" a  $0.1 \mu\text{m}$ -thick film?

- 10-3. If we irradiate PMMA with x rays of  $1.24 \text{ nm}$  wavelength, it takes about  $1000 \text{ J/cm}^3$  to fully expose the material (i.e., to get a reasonable development rate).
- Calculate the photon energy.
  - If we associate a polymer volume with each x-ray absorption event, what is that volume?
  - If we assume the volume is a cube, what would be its edge dimension?

10-4. Calculate the size of a PMMA polymer molecule of  $10^6$  molecular weight for the following assumptions:

- (a) the polymer is stretched out in a straight line;
- (b) the polymer is tightly coiled into a sphere;
- (c) the molecule is tightly coiled in a plane.

The density of PMMA is  $1.2 \text{ g/cm}^3$  and the chemical formula is  $\text{C}_5\text{H}_8\text{O}_2$ . Note that in a thin film the PMMA molecules are intertwined. Thus, they occupy a space somewhere between the two extremes (a) and (b) above.

## References

### *Problem 10-1*

Brunner, Timothy A. "Why Optical Lithography Will Live Forever."  
*J. Vac. Sci. Tech. B* 21, no. 6 (November/December 2003): 2632-2637.