

Self-Assessment: Aqueous Solutions

Weekly Homework Quiz

- (a) The value of K_a for perchloric acid, $\text{HClO}_4(aq)$, is 1×10^8 . Calculate the pH and the pOH of 1.11 M $\text{HClO}_4(aq)$ in water.
- (b) The compound, yttrium iodate, $\text{Y}(\text{IO}_3)_3$, upon dissolution in water dissociates into Y^{3+} and IO_3^- . At 37°C the solubility of $\text{Y}(\text{IO}_3)_3$ in water is 2.22×10^{-3} M. Calculate the value of the solubility product, K_{sp} , of $\text{Y}(\text{IO}_3)_3$.

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