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2.61 Internal Combustion Engines  
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MASSACHUSETTS INSTITUTE OF TECHNOLOGY  
Department of Mechanical Engineering

2.615 INTERNAL COMBUSTION ENGINES

Homework Set #3

Due: 3/4/08.

Problems:

- 1) The  $\text{NO}_x$  emission from automobiles is a mixture of NO and  $\text{NO}_2$ . At high temperatures, the mixture is mostly NO, and at low temperatures, mostly  $\text{NO}_2$ . Consider a mixture with elemental composition of 1 mole of nitrogen atom and 2 moles of oxygen atoms at a fixed pressure of 1 atmosphere. Plot the equilibrium mole fraction of NO as a function of temperature in the 600 to 1000 K range. (The actual exhaust gas is not in equilibrium; therefore, the equilibrium value of NO is a lower bound.) Note that at equilibrium above 1000 K, most of the gas is NO.

The equilibrium constants from the JANAF table are:

T(K)	$\text{Log}_{10}K_p$ for NO	$\text{Log}_{10}K_p$ for $\text{NO}_2$
600	-7.210	-6.111
700	-6.086	-5.714
800	-5.243	-5.417
900	-4.587	-5.185
1000	-4.062	-5.000

- 2) Problem 3.8 of text

(In part (c), the equilibrium constant should be  $10^{10.2}$  instead of 10.2. Also, assume that there is no dissociation of the  $\text{N}_2$  and  $\text{H}_2\text{O}$ .)

- 3) Problem 4.1 of text

Calculation of exhaust composition under fuel rich condition is a tedious exercise which is usually computerized. You can use the exhaust composition program on the web (under “good stuff”). Look at the source listing to see what is involved. If you want to, you can set  $K=3.5$  instead of 3.7; it does not make that much difference. You can also do it using the information on p. 104-105 of the text, especially table 4-3. Then you have to solve the quadratic equation 4.6 yourselves — the Matlab program on the web, exhaust\_compo, essentially does that.)

- 4) Problem 4.9 of text